

I. Multiple Choice Questions (Type-I)

1. On addition of conc. H_2SO_4 to a chloride salt, colourless fumes are evolved but in case of an iodide salt, violet fumes come out. This is because

- (i) H_2SO_4 reduces HI to I_2
- (ii) HI is of violet colour
- (iii) HI gets oxidised to I_2
- (iv) HI changes to HIO_3

Solution:

Option (iii) is the answer.

2. In qualitative analysis when H_2S is passed through an aqueous solution of salt acidified with dil. HCl , a black precipitate is obtained. On boiling the precipitate with dil. HNO_3 , it forms a solution of blue colour. Addition of excess of aqueous solution of ammonia to this solution gives _____.

- (i) a deep blue precipitate of $\text{Cu}(\text{OH})_2$
- (ii) a deep blue solution of $[\text{Cu}(\text{NH}_3)_4]^{2+}$
- (iii) a deep blue solution of $\text{Cu}(\text{NO}_3)_2$
- (iv) a deep blue solution of $\text{Cu}(\text{OH})_2 \cdot \text{Cu}(\text{NO}_3)_2$

Solution:

Option (ii) is the answer.

3. In a cyclotrimetaphosphoric acid molecule, how many single and double bonds are present?

- (i) 3 double bonds; 9 single bonds
- (ii) 6 double bonds; 6 single bonds
- (iii) 3 double bonds; 12 single bonds
- (iv) Zero double bonds; 12 single bonds

Solution;

Option (i) is the answer.

4. Which of the following elements can be involved in $p\pi-d\pi$ bonding?

- (i) Carbon
- (ii) Nitrogen
- (iii) Phosphorus
- (iv) Boron

Solution:

Option (iii) is the answer.

5. Which of the following pairs of ions are isoelectronic and isostructural?

- (i) CO_3^{2-} , NO_3^-
- (ii) ClO_3^- , CO_3^{2-}
- (iii) SO_3^{2-} , NO_3^-
- (iv) ClO_3^- , SO_3^{2-}

Solution:

Option (i) is the answer.

6. Affinity for hydrogen decreases in the group from fluorine to iodine. Which of the halogen acids should have the highest bond dissociation enthalpy?

- (i) HF
- (ii) HCl
- (iii) HBr
- (iv) HI

Solution:

Option (i) is the answer.

7. Bond dissociation enthalpy of E—H (E = element) bonds are given below. Which of the compounds will act as the strongest reducing agent?

Compound	NH ₃	PH ₃	AsH ₃	SbH ₃
$\Delta_{\text{diss}}(\text{E—H})/\text{kJ mol}^{-1}$	389	322	297	255

- (i) NH₃
- (ii) PH₃
- (iii) AsH₃
- (iv) SbH₃

Solution:

Option (iv) is the answer.

8. On heating with concentrated NaOH solution in an inert atmosphere of CO₂, white phosphorus gives a gas. Which of the following statement is incorrect about the gas?

- (i) It is highly poisonous and has smelled like rotten fish.
- (ii) It's a solution in water decomposes in the presence of light.
- (iii) It is more basic than NH₃
- (iv) It is less basic than NH₃

Solution:

Option (iii) is the answer.

9. Which of the following acids forms three series of salts?

- (i) H₃PO₂
- (ii) H₃BO₃
- (iii) H₃PO₄
- (iv) H₃PO₃

Solution:

Option (iii) is the answer.

10. Strong reducing behaviour of H₃PO₂ is due to

- (i) The low oxidation state of phosphorus
- (ii) Presence of two —OH groups and one P—H bond
- (iii) Presence of one —OH group and two P—H bonds
- (iv) High electron gain enthalpy of phosphorus

Solution:

Option (iii) is the answer.

11. On heating lead, nitrate forms oxides of nitrogen and lead. The oxides formed are _____.

- (i) N_2O , PbO
- (ii) NO_2 , PbO
- (iii) NO , PbO
- (iv) NO , PbO_2

Solution;

Option (ii) is the answer.

12. Which of the following elements does not show allotropy?

- (i) Nitrogen
- (ii) Bismuth
- (iii) Antimony
- (iv) Arsenic

Solution:

Option (i) is the answer.

13. Maximum covalency of nitrogen is _____.

- (i) 3
- (ii) 5
- (iii) 4
- (iv) 6

Solution:

Option (iii) is the answer.

14. Which of the following statements is wrong?

- (i) Single N-N bond is stronger than the single P-P bond.
- (ii) PH_3 can act as a ligand in the formation of a coordination compound with transition elements.
- (iii) NO_2 is paramagnetic.
- (iv) Covalency of nitrogen in N_2O_5 is four

Solution:

Option (i) is the answer.

15. A brown ring is formed in the ring test for NO_3^-

ion. It is due to the formation of

- (i) $[\text{Fe}(\text{H}_2\text{O})_5(\text{NO})]^{2+}$
- (ii) $\text{FeSO}_4 \cdot \text{NO}_2$
- (iii) $[\text{Fe}(\text{H}_2\text{O})_4(\text{NO})]^{2+}$
- (iv) $\text{FeSO}_4 \cdot \text{HNO}_3$

Solution:

Option (i) is the answer.

16. Elements of group-15 form compounds in +5 oxidation state. However, bismuth forms only one well-characterised compound in +5 oxidation state.

The compound is

- (i) Bi_2O_5
- (ii) BiF_5
- (iii) BiCl_5
- (iv) Bi_2S_5

Solution:

Option (ii) is the answer.

17. On heating ammonium dichromate and barium azide separately we get

- (i) N_2 in both cases
- (ii) N_2 with ammonium dichromate and NO with barium azide
- (iii) N_2O with ammonium dichromate and N_2 with barium azide
- (iv) N_2O with ammonium dichromate and NO_2 with barium azide

Solution:

Option (i) is the answer.

18. In the preparation of HNO_3 , we get NO gas by catalytic oxidation of ammonia.

The moles of NO produced by the oxidation of two moles of NH_3 will be _____.

- (i) 2
- (ii) 3
- (iii) 4
- (iv) 6

Solution:

Option (i) is the answer.

19. The oxidation state of a central atom in the anion of compound NaH_2PO_2 will be _____.

- (i) +3
- (ii) +5
- (iii) +1
- (iv) -3

Solution:

Option (iii) is the answer.

20. Which of the following is not tetrahedral in shape?

- (i) NH_4^+
- (ii) SiCl_4
- (iii) SF_4
- (iv) SO_4^{2-}

Solution:

Option (iii) is the answer.

21. Which of the following are peroxyacids of sulphur?

- (i) H_2SO_5 and $\text{H}_2\text{S}_2\text{O}_8$
- (ii) H_2SO_5 and $\text{H}_2\text{S}_2\text{O}_7$

(iii) $\text{H}_2\text{S}_2\text{O}_7$ and $\text{H}_2\text{S}_2\text{O}_8$

(iv) $\text{H}_2\text{S}_2\text{O}_6$ and $\text{H}_2\text{S}_2\text{O}_7$

Solution:

Option (i) is the answer.

22. Hot conc. H_2SO_4 acts as a moderately strong oxidising agent. It oxidises both metals and nonmetals. Which of the following element is oxidised by conc H_2SO_4 into two gaseous products?

(i) Cu

(ii) S

(iii) C

(iv) Zn

Solution:

Option (iii) is the answer.

23. A black compound of manganese reacts with a halogen acid to give greenish yellow gas. When an excess of this gas reacts with NH_3 an unstable trihalide is formed. In this process the oxidation state of nitrogen changes from _____.

(i) – 3 to +3

(ii) – 3 to 0

(iii) – 3 to +5

(iv) 0 to – 3

Solution:

Option (i) is the answer.

24. In the preparation of compounds of Xe, Bartlett had taken $\text{O}_2 + \text{Pt F}_6$ as a base compound. This is because

(i) both O_2 and Xe has the same size.

(ii) both O_2 and Xe has the same electron gain enthalpy.

(iii) both O_2 and Xe has almost same ionisation enthalpy.

(iv) both Xe and O_2 are gases.

Solution:

Option (iii) is the answer.

25. In solid state PCl_5 is a _____.

(i) covalent solid

(ii) octahedral structure

(iii) ionic solid with $[\text{PCl}_6]^+$ octahedral and $[\text{PCl}_4]^-$ tetrahedra

(iv) ionic solid with $[\text{PCl}_4]^+$ tetrahedral and $[\text{PCl}_6]^-$ octahedral

Solution:

Option (iv) is the answer.

26. Reduction potentials of some ions are given below. Arrange them in decreasing order of oxidising power.

Ion	ClO_4^-	IO_4^-	BrO_4^-
Reduction potential E°/V	$E^\circ = 1.19\text{V}$	$E^\circ = 1.65\text{V}$	$E^\circ = 1.74\text{V}$

- (i) $\text{ClO}_4^- \rightarrow \text{IO}_4^- \rightarrow \text{BrO}_4^-$
(ii) $\text{IO}_4^- \rightarrow \text{BrO}_4^- \rightarrow \text{ClO}_4^-$
(iii) $\text{BrO}_4^- \rightarrow \text{IO}_4^- \rightarrow \text{ClO}_4^-$
(iv) $\text{BrO}_4^- \rightarrow \text{ClO}_4^- \rightarrow \text{IO}_4^-$

Solution:

Option (iii) is the answer.

27. Which of the following is isoelectronic pair?

- (i) ICl_2 , ClO_2
(ii) BrO_2^- , BrF_2^+
(iii) ClO_2 , BrF
(iv) CN^- , O_3

Solution:

Option (ii) is the answer.

II. Multiple Choice Questions (Type-II)

Note: In the following questions two or more options may be correct.

28. If chlorine gas is passed through a hot NaOH solution, two changes are observed in the oxidation number of chlorine during the reaction. These are _____ and _____.

- (i) 0 to +5
(ii) 0 to +3
(iii) 0 to -1
(iv) 0 to +1

Solution:

Option (i) and (iii) are the answers.

29. Which of the following options are not in accordance with the property mentioned against them?

- (i) $\text{F}_2 > \text{Cl}_2 > \text{Br}_2 > \text{I}_2$
(ii) $\text{MI} > \text{MBr} > \text{MCl} > \text{MF}$
(iii) $\text{F}_2 > \text{Cl}_2 > \text{Br}_2 > \text{I}_2$
(iv) $\text{HI} < \text{HBr} < \text{HCl} < \text{HF}$

Oxidising power.

Ionic character of metal halide.

Bond dissociation enthalpy.

Hydrogen-halogen bond strength.

Solution:

Option (ii) and (iii) are the answers

30. Which of the following is correct for P_4 the molecule of white phosphorus?

- (i) It has 6 lone pairs of electrons.
(ii) It has six P-P single bonds.
(iii) It has three P-P single bonds.
(iv) It has four lone pairs of electrons.

Solution:

Option (ii) and (iv) are the answers.

31. Which of the following statements are correct?

- (i) Among halogens, radius ratio between iodine and fluorine is maximum.
- (ii) Leaving F—F bond, all halogens have weaker X—X bond than X—X' bond in interhalogens.
- (iii) Among interhalogen compounds, the maximum number of atoms are present in iodine fluoride.
- (iv) Interhalogen compounds are more reactive than halogen compounds.

Solution:

Option (i), (iii) and (iv) are the answers.

32. Which of the following statements are correct for SO₂ gas?

- (i) It acts as a bleaching agent in moist conditions.
- (ii) Its molecule has linear geometry.
- (iii) It's dilute solution is used as a disinfectant.
- (iv) It can be prepared by the reaction of dilute H₂SO₄ with metal sulphide.

Solution:

Option (i) and (iii) are the answers.

33. Which of the following statements are correct?

- (i) All three N—O bond lengths in HNO₃ are equal.
- (ii) All P—Cl bond lengths in PCl₅ molecule in the gaseous state are equal.
- (iii) P₄ molecule in white phosphorus have angular strain therefore white phosphorus is very reactive.
- (iv) PCl is ionic in the solid-state in which cation is tetrahedral and the anion is octahedral.

Solution:

Option (iii) and (iv) are the answers.

34. Which of the following orders are correct as per the properties mentioned against each?

- | | |
|---|---------------------------------------|
| (i) As ₂ O ₃ < SiO ₂ < P ₂ O ₃ < SO ₂ | Acid strength. |
| (ii) AsH ₃ < PH ₃ < NH ₃ | Enthalpy of vaporization. |
| (iii) S < O < Cl < F | More negative electron gain enthalpy. |
| (iv) H ₂ O > H ₂ S > H ₂ Se > H ₂ Te | Thermal stability. |

Solution:

Option (i) and (iv) are the answers.

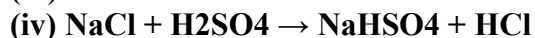
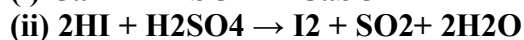
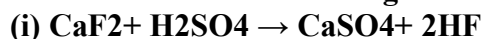
35. Which of the following statements are correct?

- (i) S—S bond is present in H₂S₂O₆
- (ii) In peroxosulphuric acid (H₂SO₅) sulphur is in +6 oxidation state.
- (iii) Iron powder along with Al₂O₃ and K₂O is used as a catalyst in the preparation of NH₃ by Haber's process.
- (iv) Change in enthalpy is positive for the preparation of SO₃ by catalytic oxidation of SO₂

Solution:

Option (i) and (ii) are the answers.

36. In which of the following reactions conc. H_2SO_4 is used as an oxidising reagent?



Solution:

Option (ii) and (iii) are the answers.

37. Which of the following statements are true?

(i) The only type of interactions between particles of noble gases is due to weak dispersion forces.

(ii) Ionisation enthalpy of molecular oxygen is very close to that of xenon.

(iii) Hydrolysis of XeF_6 is a redox reaction.

(iv) Xenon fluorides are not reactive.

Solution:

Option (i) and (ii) are the answers.

III. Short Answer Type

38. In the preparation of H_2SO_4 by Contact Process, why is SO_3 not absorbed directly in water to form H_2SO_4 ?

Solution:

Dissolution of SO_3 in water is highly exothermic. This leads to the formation of a mist of tiny droplets which are highly corrosive and they can even attack the lead pipelines (lead pipelines are used to cover the tower in contact process).

39. Write a balanced chemical equation for the reaction showing catalytic oxidation of NH_3 by atmospheric oxygen.

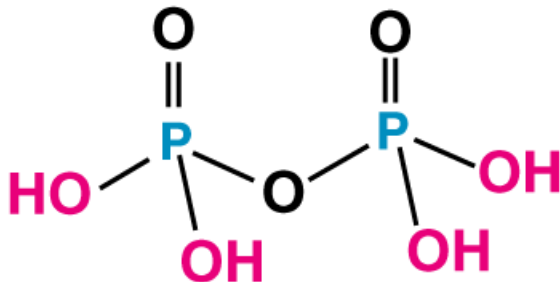
Solution:



Pt gauze reacts as a catalyst, which is added to increase the rate of reaction.

40. Write the structure of pyrophosphoric acid.

Solution:



41. PH₃ forms bubbles when passed slowly in water but NH₃ dissolves. Explain why?

Solution:

PH₃ does not form hydrogen bonding with water. Therefore it is not soluble in water and it escapes as a gas and forms bubbles whereas ammonia forms hydrogen bonding and is soluble in water.

42. In PCl₅, phosphorus is in sp³d hybridised state but all its five bonds are not equivalent. Justify your answer with reason.

Solution:

The size of axial bonds is greater than the size of equatorial bonds to overcome repulsion because the three equatorial bonds cause more repulsion. Therefore two axial P-Cl bonds are longer and different from equatorial bonds.

43. Why is nitric oxide paramagnetic in gaseous state but the solid obtained on cooling it is diamagnetic?

Solution:

Nitric acid in the gaseous state exists in monomer form. It consists of only one unpaired electron therefore it is paramagnetic. In solid-state, it exists as a dimer (N₂O₂). There is no unpaired electron in its dimer form, therefore it is diamagnetic.

44. Give the reason to explain why ClF₃ exists but FCl₃ does not exist.

Solution:

Chlorine has vacant d orbitals hence it can show an oxidation state of +3. Fluorine has no d orbitals, it cannot show a positive oxidation state. Fluorine shows only -1 oxidation state. Therefore FCl₃ does not exist.

45. Out of H₂O and H₂S, which one has a higher bond angle and why?

Solution:

H₂O has higher bond angle than H₂S because as we move from oxygen to sulphur the size of the central atom increases and electronegativity decreases due to which bond pair goes away from the central atom which results in a decrease in bond pair repulsion and hence bond angle decreases.

46. SF₆ is known but SCl₆ is not. Why?

Solution:

Fluorine is the strongest oxidizing agent and it can oxidize sulphur to its maximum oxidation state +6 to

form SF₆. Chlorine is not a good oxidizing agent, it cannot oxidize sulphur to its maximum oxidation state. Chlorine can oxidize sulphur to only +4 oxidation state. Hence it can form SCl₄ but not SCl₆.

47. On reaction with Cl₂, phosphorus forms two types of halides 'A' and 'B'. Halide A is a yellowish-white powder but halide 'B' is a colourless oily liquid. Identify A and B and write the formulas of their hydrolysis products.

Solution:

Halide A is PCl₅ because it is a yellowish-white powder

Halide B is PCl₃ because it is a colourless oily liquid.

PCl₃ + 3H₂O → H₃PO₃ + 3HCl

PCl₅ undergoes a violent hydrolysis

PCl₅ + H₂O → POCl₃ + 2HCl

PCl₅ + 4H₂O → H₃PO₄ + 5HCl

48. In the ring test of NO₃⁻ ion, Fe²⁺ ion reduces nitrate ion to nitric oxide, which combines with Fe²⁺ (aq) ion to form the brown complex. Write the reactions involved in the formation of a brown ring.

Solution:

NO₃⁻ + 3Fe²⁺ + 4H⁺ → NO + 3Fe³⁺ + 2H₂O

[Fe(H₂O)₆]SO₄ + NO → [Fe(H₂O)₅NO]SO₄

Thus the complex formed is brown.

49. Explain why the stability of oxoacids of chlorine increases in the order given below:

HClO < HClO₂ < HClO₃ < HClO₄

Solution:

As the electronegativity of halogen decreases, the tendency of XO₃ group (X = halogens) to withdraw electrons of the O-H bond towards itself decreases and hence the acid strength of the perhalic acid decreases.

50. Explain why ozone is thermodynamically less stable than oxygen.

Solution:

Ozone is thermodynamically less stable because it decomposes into oxygen and this decomposition leads result in the liberation of heat, so its entropy is positive and free energy is negative.

51. P₄O₆ reacts with water according to equation P₄O₆ + 6H₂O → 4H₃PO₃.

Calculate the volume of 0.1 M NaOH solution required to neutralise the acid formed by dissolving 1.1 g of P₄O₆ in H₂O.

Solution:

P₄O₆ + 6H₂O → 4H₃PO₃

4H₃PO₃ + 8NaOH → 4Na₂HPO₃ + 8H₂O

Overall reaction :

P₄O₆ + 8NaOH → 4Na₂HPO₃ + 2H₂O

Moles of P₄O₆ = 1.1/220 = 0.005

Acid formed by one mole of P₄O₆ requires = 8 mol

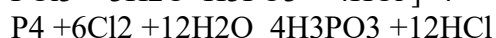
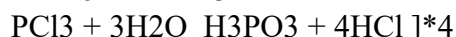
Acid formed by 0.005 mol of P₄O₆ requires = 0.04 mol NaOH is present in 100mL solution

0.04 mol of NaOH is present in solution = 1000/0.1 * 0.04 = 400mL

52. White phosphorus reacts with chlorine and the product hydrolyses in the presence of water. Calculate the mass of HCl obtained by the hydrolysis of the product formed by the reaction of 62 g of white phosphorus with chlorine in the presence of water.

Solution:

When white phosphorous reacts with chlorine:



Moles of white P = $62/124 = 0.5 \text{ mol}$

1 mol of white P₄ produces HCl = 12 mol

0.5 mol of white P₄ will produce HCl = $12 \times 0.5 = 6 \text{ mol}$

Mass of HCl = $6 \times 36.5 = 219.0 \text{ g}$

53. Name three oxoacids of nitrogen. Write the disproportionation reaction of that oxoacid of nitrogen in which nitrogen is in +3 oxidation state.

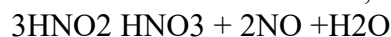
Solution:

There are three oxoacids of nitrogen.

Nitric acid (HNO₃), Nitrous acid (HNO₂), Hyponitrous acid (H₂N₂O₂)

+3 oxidation is shown by HNO₂, Therefore it undergoes disproportion reaction,

To calculate oxidation state, consider the oxidation state of N is x.



Oxidation state of nitrogen in HNO₂ is +3

$$X + (+1) + 2 \times (-2) = 0, x = +3$$

Oxidation state of nitrogen in HNO₃ is +5

$$X + (+1) + 3 \times (-2) = 0, x = 5$$

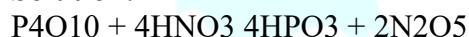
Oxidation state of nitrogen in NO is +2

$$X + (-2) = 0, x = 2$$

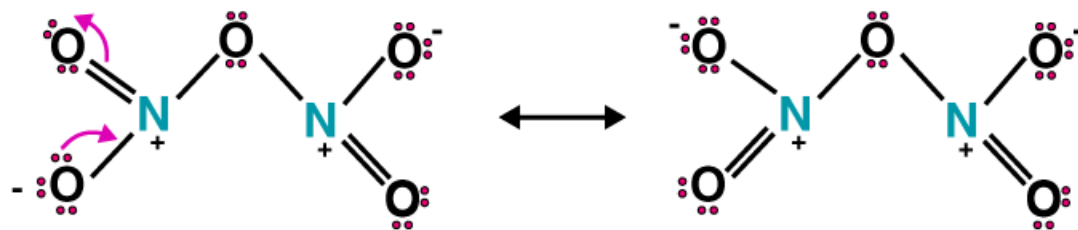
Therefore +3 oxidation state changes to +5 and +2 oxidation states.

54. Nitric acid forms an oxide of nitrogen on reaction with P₄O₁₀. Write the reaction involved. Also, write the resonating structures of the oxide of nitrogen formed.

Solution:



White phosphorous is very reactive as compared to red phosphorous due to angular strain in white phosphorous.



55. Phosphorus has three allotropic forms — (i) white phosphorus (ii) red phosphorus and (iii) black phosphorus. Write the difference between white and red phosphorus on the basis of their structure and reactivity.

Solution:

1. White phosphorus

(i) It is soft waxy solid having garlic order

(ii) It is poisonous

(iii) Has low melting point and boiling point because P₄ molecules are held together by weak van der Waal forces of attraction

2. Red Phosphorus

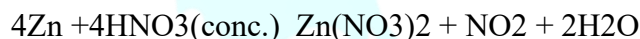
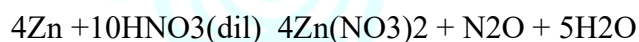
(i) It is hard, crystalline, odourless solid

(ii) It is non-poisonous

(iii) It has a high melting point because of its polymeric structure.

56. Give an example to show the effect of concentration of nitric acid on the formation of oxidation product.

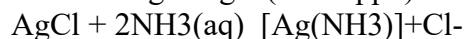
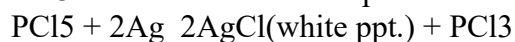
Solution:



57. PCI₅ reacts with finely divided silver on heating and a white silver salt is obtained, which dissolves on adding excess aqueous NH₃ solution. Write the reactions involved to explain what happens.

Solution:

PCI₅ reacts with silver to form white silver salt(AgCl). Which than dissolves on adding excess aqueous NH₃ to form a soluble complex



58. Phosphorus forms a number of oxoacids. Out of these oxoacidsphosphinicacid has strong

reducing property. Write its structure and also write a reaction showing its reducing behaviour.

Solution:

